

What You Should Know from “the Bohr atom”!

- 1.) What is the Bohr atom?
- 2.) How does the Bohr atom fall short as an accurate model?
- 3.) Where did the idea of energy levels come from?
- 4.) Energy levels are measured in NEGATIVE electron-volts? What is the negative sign for? That is, what does a negative energy tell you?
- 5.) What is the ground level?
- 6.) What is the ground state energy value for hydrogen?
- 7.) What is a spectroscope?

- 8.) What information will a spectroscope help you to calculate that will help determine an atom's energy levels values?
- 9.) Be able to determine the number of different frequencies of E/M wave given off due to a given number of energy levels.
- 10.) What are the relationships you need to use to start with a wavelength determined by a spectroscope and end with the size of the energy level jumps in an atom?
- 11.) What is the first excited state in an atom?
- 12.) The group of hydrogen lines that terminate at the ground level are called what? Are they in the optical range?
- 13.) The group of hydrogen lines that terminate at the first excited state are called what? Are they in the optical range?